

CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/01

Paper 1 Multiple Choice

May/June 2003

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the answer sheet in the spaces provided unless this has been done for you.

There are forty questions on this paper. Answer **all** questions. For each question, there are four possible answers **A, B, C, and D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate answer sheet.

Read the instructions on the Answer Sheet very carefully.

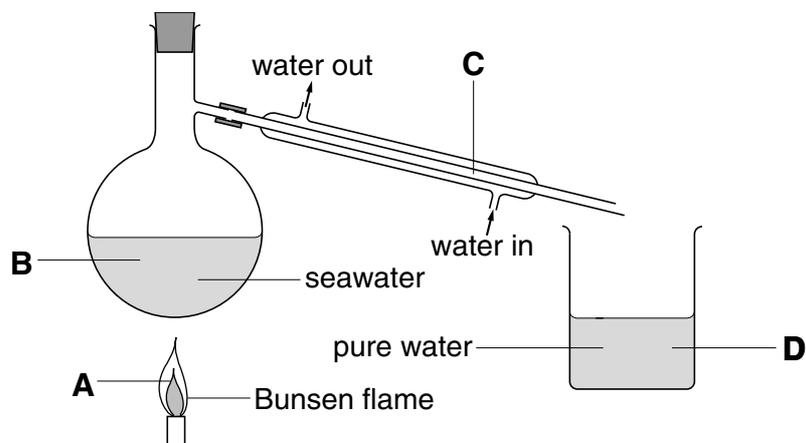
Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

- 1 The diagram shows how to obtain pure water from seawater.

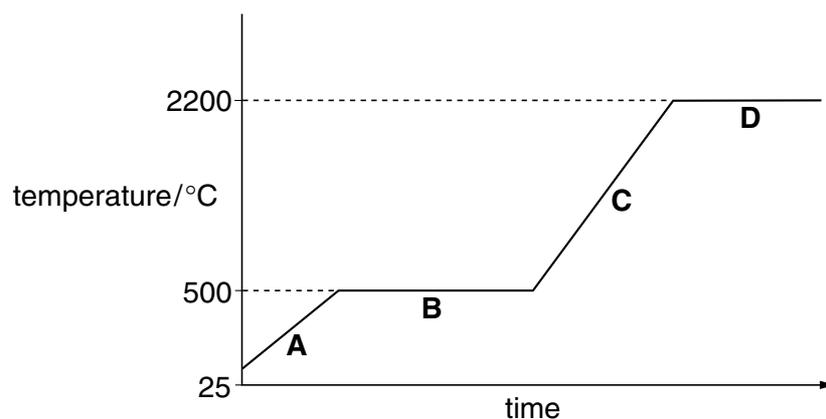
Where do water molecules lose energy?



- 2 A solid metal is heated until it turns to vapour.

The graph shows the temperature of the metal during this process.

Which part of the graph shows the melting of the metal?



- 3 Some chemical compounds are purified by recrystallisation.

What can be used to test the purity of the crystals?

- A melting point
- B colour of crystals
- C size of crystals
- D solubility

- 4 What could be the melting point and boiling point of water containing a dissolved impurity?

	melting point / °C	boiling point / °C
A	+3	96
B	+3	104
C	-3	96
D	-3	104

- 5 Which number in the table is -1?

particle	charge	relative mass
electron	A	B
neutron	C	1
proton	D	1

- 6 What is the electronic structure of an atom with a proton number 5 and a nucleon number 11?

A 1, 8, 2 **B** 2, 8, 1 **C** 2, 3 **D** 3, 2

- 7 What changes when an ion is made from an atom?

- A** the number of electrons only
B the number of neutrons only
C the number of protons only
D the number both of protons and of neutrons

- 8 Strontium, Sr, is a metal that forms an ionic chloride SrCl_2 .

Sulphur, S, is a non-metal that forms a covalent chloride SCl_2 .

Which compound is likely to have the higher melting point (m.p.) and which is more soluble in water?

	higher m.p.	more soluble in water
A	SrCl_2	SrCl_2
B	SrCl_2	SCl_2
C	SCl_2	SrCl_2
D	SCl_2	SCl_2

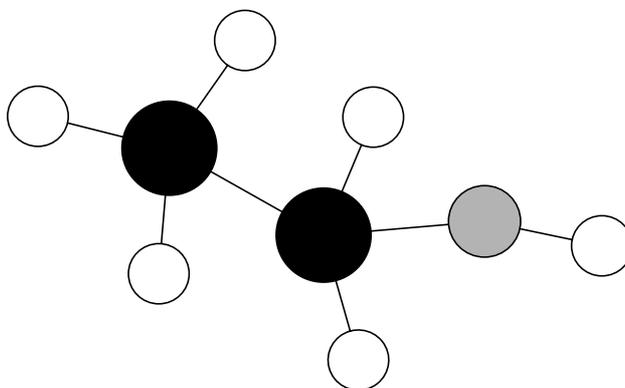
- 9 The relative atomic mass of oxygen is 16 and that of hydrogen is 1.

This means that ... (i) ... of oxygen has the same mass as ... (ii) ... of hydrogen.

Which words correctly complete the gaps?

	gap (i)	gap (ii)
A	an atom	thirty-two molecules
B	an atom	eight molecules
C	a molecule	sixteen atoms
D	a molecule	eight atoms

- 10 The diagram shows a model of a molecule containing carbon, hydrogen and oxygen.



How many atoms of each element are in the molecule?

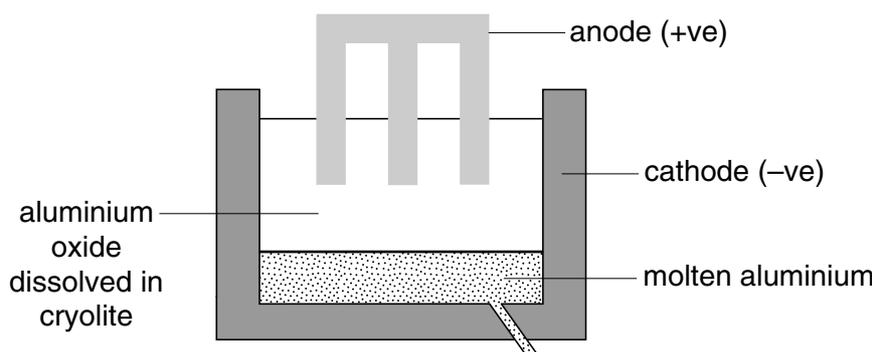
	carbon	hydrogen	oxygen
A	1	6	2
B	2	5	1
C	2	6	1
D	6	2	1

- 11 Water is formed when 48 g of oxygen combine with 6 g of hydrogen.

What mass of oxygen combines with 2 g of hydrogen?

- A** 12 g **B** 16 g **C** 96 g **D** 144 g

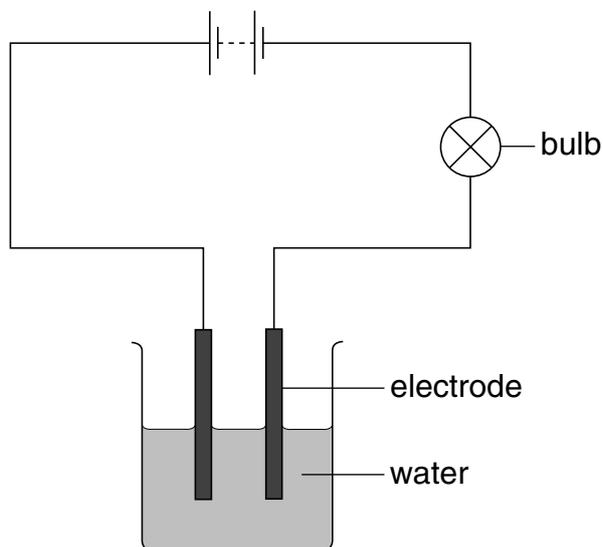
12 The diagram shows how aluminium is manufactured by electrolysis.



What are the anode and cathode made of?

	anode	cathode
A	aluminium	aluminium
B	aluminium	graphite
C	graphite	aluminium
D	graphite	graphite

13 A student sets up the apparatus shown. The bulb does not light.



After the student adds substance **X** to the water, the bulb lights.

What is **X**?

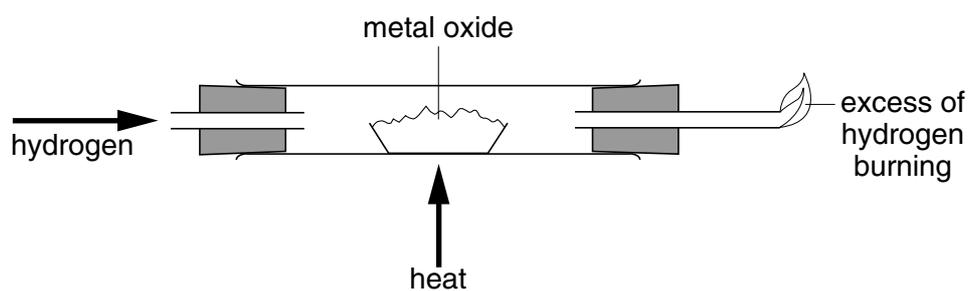
- A** calcium carbonate
- B** carbon
- C** copper(II) sulphate
- D** ethanol

14 The following elements have radioactive isotopes.

Which element is used as a source of energy because of its radioactivity?

- A carbon
- B hydrogen
- C iodine
- D uranium

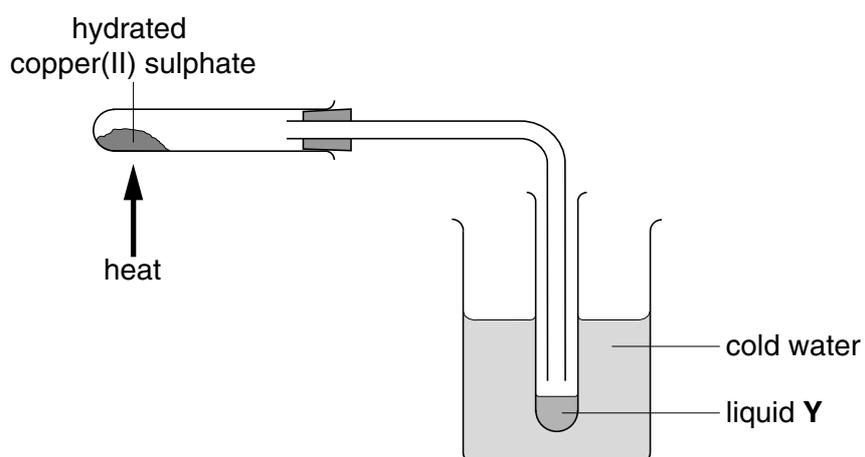
15 When hydrogen is passed over a heated metal oxide, the metal and steam are formed.



What happens to the hydrogen and to the metal oxide?

	hydrogen	metal oxide
A	oxidised	oxidised
B	oxidised	reduced
C	reduced	oxidised
D	reduced	reduced

- 16 When hydrated copper(II) sulphate is heated in the apparatus shown, solid X and liquid Y are produced.



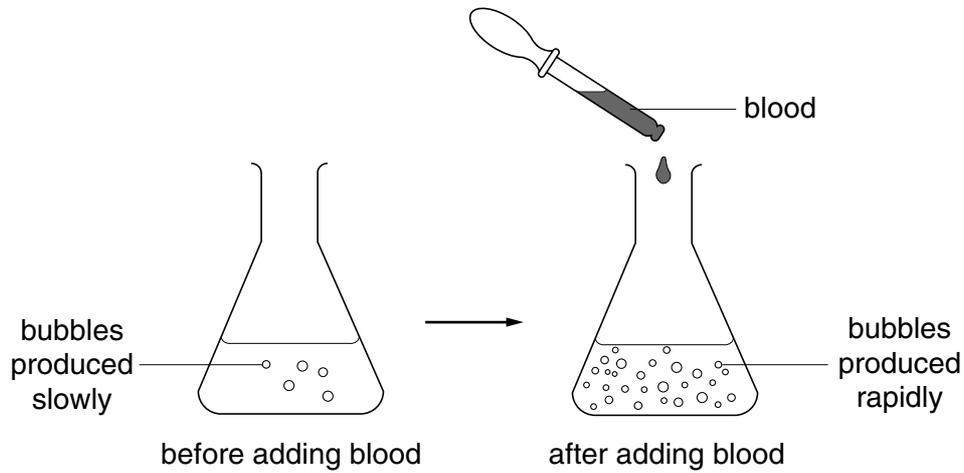
Which changes are noticed when liquid Y is added to cold solid X?

	colour change	heat change
A	blue to white	heat given out
B	blue to white	heat taken in
C	white to blue	heat given out
D	white to blue	heat taken in

17 A solution of hydrogen peroxide releases oxygen slowly at room temperature.



The diagrams show the effect of adding blood to the solution.

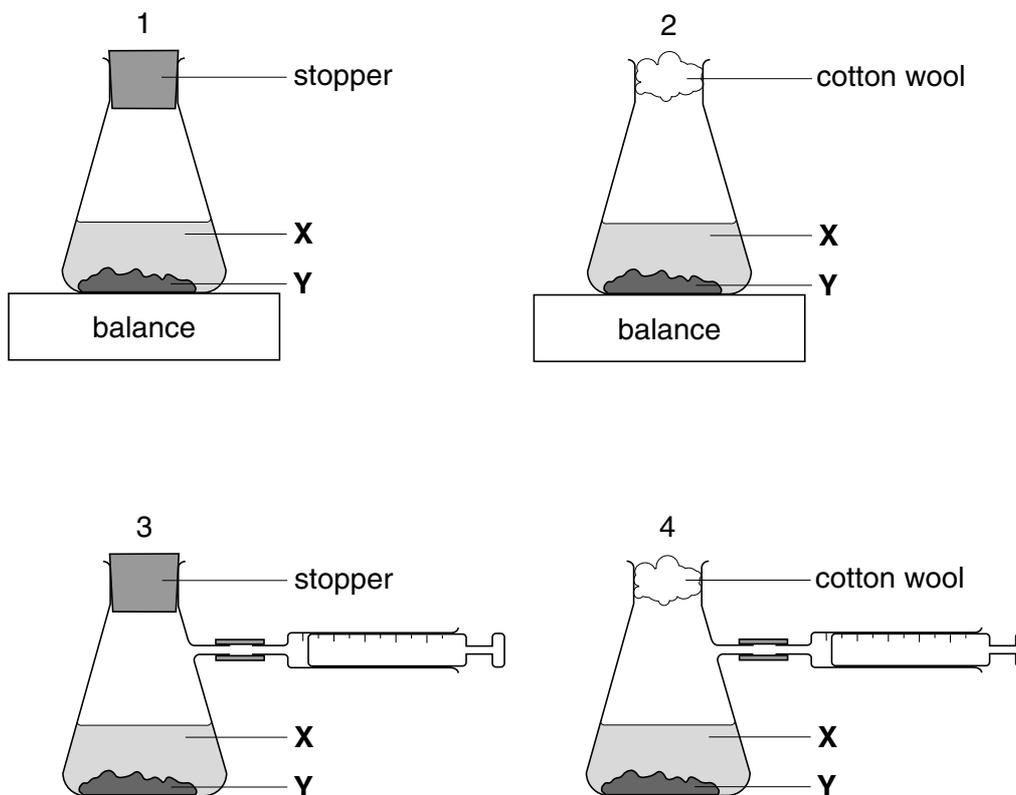


What could be the reason for the observed change?

- A Blood contains an enzyme.
- B Blood contains water.
- C The hydrogen peroxide becomes more concentrated.
- D The hydrogen peroxide is neutralised by blood.

18 A liquid **X** reacts with solid **Y** to form a gas.

Which two diagrams show suitable methods for investigating the speed of the reaction?



- A 1 and 3
- B 1 and 4
- C 2 and 3
- D 2 and 4

19 Which substance does **not** form copper(II) sulphate with warm, dilute sulphuric acid?

- A copper
- B copper(II) carbonate
- C copper(II) hydroxide
- D copper(II) oxide

20 Which test method and gas are correctly linked?

	test method	gas
A	a lighted splint	oxygen
B	a glowing splint	hydrogen
C	damp litmus paper	chlorine
D	limewater	ammonia

21 Water is added to a test-tube containing dilute sulphuric acid of pH 4.

What could be the pH of the resulting solution?

- A** 8 **B** 6 **C** 4 **D** 2

22 Magnesium, on the left of Period Two of the Periodic Table, is more metallic than chlorine on the right of this Period.

Why is this?

Magnesium has

- A** fewer electrons.
B fewer protons.
C fewer full shells of electrons.
D fewer outermost electrons.

23 An inert gas **X** is used to fill weather balloons.

Which descriptions of **X** are correct?

	number of outer electrons in atoms of X	structure of gas X
A	2	single atoms
B	2	diatomic molecules
C	8	single atoms
D	8	diatomic molecules

24 A student is asked to complete two sentences.

Metallic and non-metallic elements are classified in the ... (i) ... This can be used to ... (ii) ... properties of elements.

Which words correctly complete the gaps?

	gap (i)	gap (ii)
A	Periodic Table	measure
B	Periodic Table	predict
C	reactivity series	measure
D	reactivity series	predict

25 Which material is an alloy that contains a non-metallic element?

- A** brass
- B** haematite
- C** manganese
- D** steel

26 The table gives information about the reactivity of three metals P, Q and R.

metal	reaction with air	reaction with steam	reaction with dilute hydrochloric acid
P	burns with sparks	forms an oxide	forms hydrogen
Q	slowly forms an oxide	no reaction	no reaction
R	slowly forms an oxide	no reaction	forms hydrogen

What is the order of reactivity of P, Q and R?

	most reactive	—————→	least reactive
A	P	Q	R
B	P	R	Q
C	Q	R	P
D	R	Q	P

27 The bodies of aircraft are often made using aluminium.

Which **two** properties of aluminium make it suitable for this purpose?

	property 1	property 2
A	good conductor of electricity	good conductor of heat
B	good conductor of electricity	strong
C	good conductor of heat	low density
D	strong	low density

28 Which raw materials are used in the manufacture of iron?

- A** bauxite and lime
- B** bauxite and limestone
- C** haematite and lime
- D** haematite and limestone

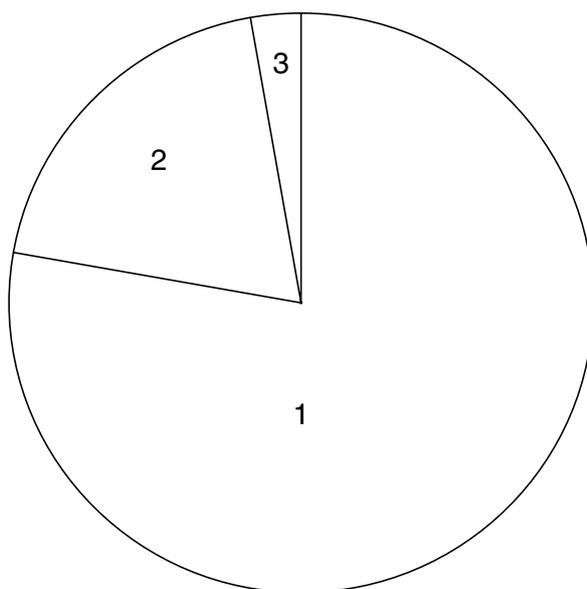
29 In a car industry, approximately 45 000 litres of water are required to produce a single car.

This water does not need to be very pure.

Which purification methods would be suitable and economic to use?

	chlorinated	distilled
A	✓	✓
B	✓	✗
C	✗	✓
D	✗	✗

30 The pie-chart shows the composition of air.



What are the gases in parts 1, 2 and 3 of the pie-chart?

	1	2	3
A	nitrogen	other gases	oxygen
B	nitrogen	oxygen	other gases
C	oxygen	other gases	nitrogen
D	oxygen	nitrogen	other gases

31 A steel works and a chemical works are built near to a city. The limestone buildings in the city begin to crumble.

Which gas is most likely to cause this damage?

- A** carbon dioxide
- B** carbon monoxide
- C** oxygen
- D** sulphur dioxide

32 Which methods can be used to prevent the rusting of an iron girder of a bridge?

	coat it with grease	electroplate it	paint it
A	✓	✓	✓
B	✓	✓	✗
C	✗	✓	✓
D	✗	✗	✓

33 A student heats a mixture of ammonium chloride and calcium hydroxide. She tests the gas given off with damp red litmus paper.

What is the name of the gas and the final colour of the litmus paper?

	gas	colour
A	ammonia	blue
B	ammonia	red
C	chlorine	red
D	chlorine	white

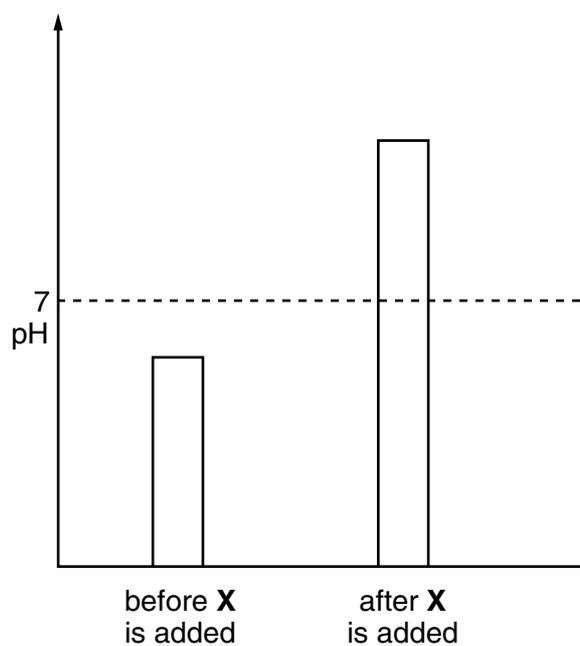
34 A newspaper article claims that carbon dioxide is formed as follows.

- 1 during respiration
- 2 when calcium carbonate reacts with hydrochloric acid
- 3 when methane burns in air

Which statements are correct?

- A** 1, 2 and 3
B 1 and 2 only
C 1 and 3 only
D 2 and 3 only

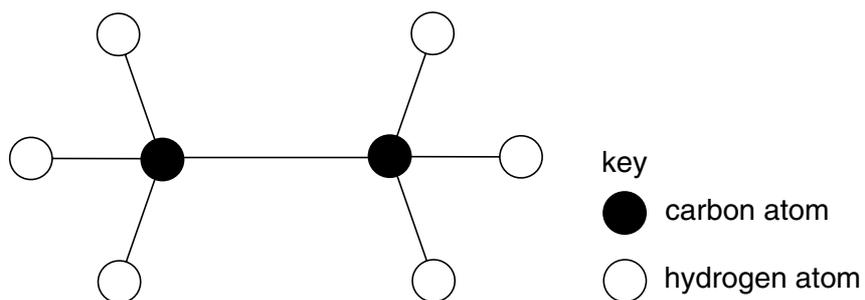
35 The diagram shows how the pH of an industrial waste changes when substance X is



What is substance X?

- A coal
- B lime
- C salt
- D water

36 The diagram shows a model of an organic compound.



What is the name of this compound?

- A ethane
- B ethanoic acid
- C ethanol
- D ethene

37 Bitumen is a substance obtained from the fractional distillation of petroleum.

What are the boiling points and the sizes of the molecules in bitumen?

	boiling points	sizes of molecules
A	high	large
B	high	small
C	low	large
D	low	small

38 Which hydrocarbons in the table are members of the same homologous series?

hydrocarbon	1	2	3	4
state at room temperature	gas	gas	liquid	liquid
reaction with oxygen	burns	burns	burns	burns
aqueous reaction with bromine	decolourises bromine	no reaction	decolourises bromine	no reaction

- A** 1 and 2
- B** 1 and 3
- C** 3 and 4
- D** 1, 2, 3 and 4

DATA SHEET
The Periodic Table of the Elements

		Group													
I	II	III	IV	V	VI	VII	0								
7 Li Lithium	9 Be Beryllium	1 H Hydrogen										4 He Helium			
23 Na Sodium	24 Mg Magnesium	11 B Boron	12 C Carbon	14 N Nitrogen	16 O Oxygen	19 F Fluorine	20 Ne Neon					10			
39 K Potassium	40 Ca Calcium	27 Al Aluminium	28 Si Silicon	31 P Phosphorus	32 S Sulphur	35.5 Cl Chlorine	40 Ar Argon					18			
85 Rb Rubidium	88 Sr Strontium	65 Zn Zinc	64 Cu Copper	59 Ni Nickel	56 Fe Iron	55 Mn Manganese	58 Co Cobalt	70 Ga Gallium	73 Ge Germanium	75 As Arsenic	79 Se Selenium	84 Kr Krypton			
133 Cs Caesium	137 Ba Barium	112 Cd Cadmium	108 Ag Silver	106 Pd Palladium	101 Ru Ruthenium	103 Rh Rhodium	103 Rh Rhodium	115 In Indium	119 Sn Tin	122 Sb Antimony	128 Te Tellurium	131 Xe Xenon			
226 Fr Francium	227 Ra Radium	201 Hg Mercury	197 Au Gold	195 Pt Platinum	190 Os Osmium	186 Re Rhenium	192 Ir Iridium	204 Tl Thallium	207 Pb Lead	209 Bi Bismuth	210 Po Polonium	222 Rn Radon			
<p>3-71 Lanthanoid series 0-103 Actinoid series</p> <div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 5px;"> <table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 20px; text-align: center;">a</td> <td style="width: 20px; text-align: center;">X</td> <td style="width: 20px; text-align: center;">b</td> </tr> </table> </div> <div style="font-size: 0.8em;"> <p>a = relative atomic mass X = atomic symbol b = proton (atomic) number</p> </div> </div>													a	X	b
a	X	b													
140 Ce Cerium	141 Pr Praseodymium	144 Nd Neodymium	150 Sm Samarium	152 Eu Europium	157 Gd Gadolinium	162 Dy Dysprosium	165 Ho Holmium	167 Er Erbium	169 Tm Thulium	173 Yb Ytterbium	175 Lu Lutetium	232 Th Thorium			
58	59	60	62	63	64	66	67	68	69	70	71	90			
238 U Uranium	238 Pa Protactinium	238 U Uranium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	106 Lr Lawrencium	86			

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).